

## Chemistry Chapter 12 Solution Manual Stoichiometry

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### Chemistry Chapter 12 Solution Manual

CHAPTER 12 SOLUTIONS MANUAL States of Matter Section 12.1 Gases pages 402–410 Practice Problems page 405 1. Calculate the ratio of effusion rates for nitrogen (N<sub>2</sub>) and neon (Ne).  $\frac{\text{Rate N}_2}{\text{Rate Ne}} = 0.849$   $\frac{\text{Rate N}_2}{\text{Rate Ne}} = \sqrt{\frac{\text{molar mass Ne}}{\text{molar mass N}_2}}$   $\sqrt{\frac{20.18}{28.02}} = 0.849$  2. Calculate the ratio of diffusion rates for carbon monoxide and carbon dioxide.  $\frac{\text{Rate CO}}{\text{Rate CO}_2} = \sqrt{\frac{\text{molar mass CO}_2}{\text{molar mass CO}}}$

### States of Matter

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4.2 For the reaction,  $2A + B \longrightarrow A_2B$ , the rate =  $k[A][B]^2$  with  $k = 2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$ . Calculate the initial rate of the reaction when  $[A] = 0.1 \text{ mol L}^{-1}$ ,  $[B] = 0.2 \text{ mol L}^{-1}$ . Calculate the rate of reaction after  $[A]$  is reduced to  $0.06 \text{ mol L}^{-1}$ . Sol. Initial rate of reaction =  $k[A][B]^2 = (2.0 \times 10^{-6} \text{ mol}^{-2} \text{ s}^{-1})(0.1 \text{ mol L}^{-1})(0.2 \text{ mol L}^{-1})^2 = 8 \times 10^{-9} \text{ mol L}^{-1} \text{ s}^{-1}$ .

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